

Surname	Centre Number	Candidate Number
Other Names		0



GCSE

3440UB0-1



**APPLIED SCIENCE (Single Award)
UNIT 2: Science to Support our Lifestyles**

HIGHER TIER

TUESDAY, 14 MAY 2019 – AFTERNOON

1 hour 30 minutes

For Examiner's use only		
Question	Maximum Mark	Mark Awarded
1.	19	
2.	6	
3.	8	
4.	6	
5.	10	
6.	8	
7.	9	
8.	9	
Total	75	

ADDITIONAL MATERIALS

A calculator.

INSTRUCTIONS TO CANDIDATES

Use black ink or black ball-point pen.

Write your name, centre number and candidate number in the spaces at the top of this page.

Answer **all** questions.

Write your answers in the spaces provided in this booklet.

INFORMATION FOR CANDIDATES

The number of marks is given in brackets at the end of each question or part-question.

Question 4 is a quality of extended response (QER) question where your writing skills will be assessed.

The Periodic Table is printed on page 20 of this examination paper.

Answer **all** questions.

1. Food manufacturers are required to measure the amount of energy contained in their food products.

(a) A cereal bar has the following information on its label.

	Per 100 g	Per cereal bar	RDA
Energy (kJ)	1966	787	8400

- (i) Calculate the percentage of the recommended daily allowance (RDA) provided by **one** cereal bar. Give your answer to 2 significant figures. [2]

Percentage = %

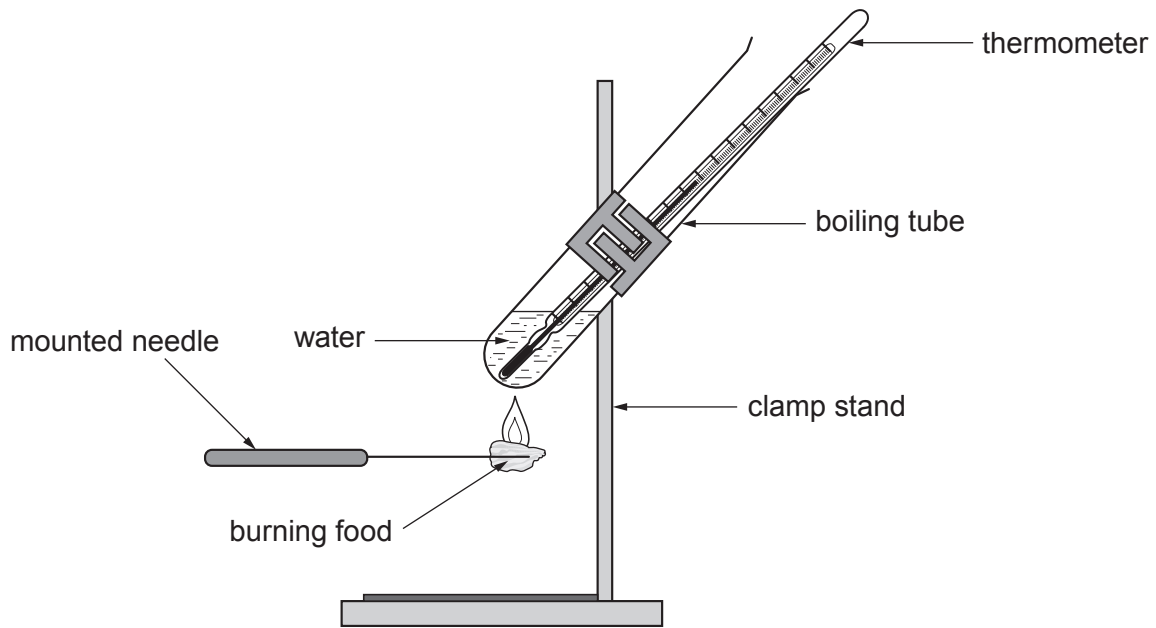
- (ii) Calculate the mass in grams of **one** cereal bar. [2]

Mass = g

- (iii) Give a reason for not using the value for energy **per cereal bar** when comparing different brands. [1]

- (iv) Explain why the Government is concerned about constantly exceeding the RDA for energy. [3]
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(b) A student used the following equipment to find the energy content of some foods.



(i) State the independent variable in this experiment. [1]

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(ii) State the dependent variable in this experiment. [1]

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(iii) State **one** variable that should be controlled in this experiment. [1]

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(c) The student obtained the following results.

Type of food	Mass of water (g)	Mass of food burned (g)	Temperature at start (°C)	Temperature at end (°C)	Temperature increase (°C)	Energy released (J)	Energy released per gram (J/g)
cheese biscuit	20	3.0	20	56	36	2268	756
corn snack	20	0.5	21	36	15	1848
digestive biscuit	20	4.0	20	93	73	1636	409
cereal bar	20	4.0	22	48			

(i) How would the student make sure the results are reproducible? [1]

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(ii) Calculate the energy released by the **corn snack**. Write this value in the table. [1]
Space for working

(iii) Use the information in the table and the equation:

$$\text{Energy released per gram (J/g)} = \frac{\text{mass of water (g)} \times \text{temperature increase (°C)} \times 4.2}{\text{mass of food sample (g)}}$$

to calculate the energy released per gram for the **cereal bar**. [3]

Energy released per gram = J/g

- (iv) The student suggests that a 30 g packet of corn snacks contains half the energy of four digestive biscuits of total mass 60 g.

Explain whether the student is correct.
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[3]

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2. Rugby players are prone to knee injuries, some of which lead to long-term pain through osteoarthritis.

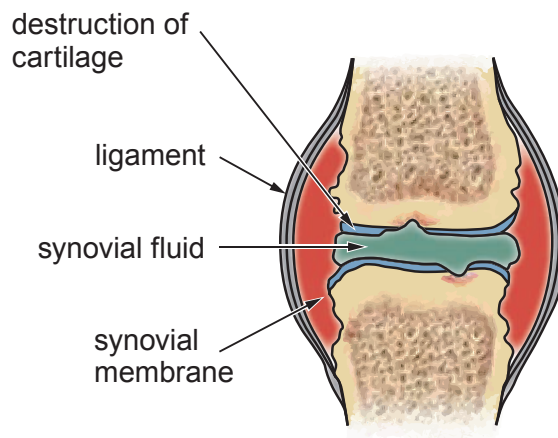
(a) Describe how the synovial membrane helps to maintain a healthy knee joint. [2]

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(b) The diagram below shows a knee joint damaged by osteoarthritis.



Explain what has happened to cause pain in the joint. [4]

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3. Cardiff University has developed a drug called INX-189 that treats the liver disease, hepatitis C. The World Health Organisation estimates 170 million people are carriers of the hepatitis C virus worldwide. More than 350 000 people die from hepatitis-related illnesses every year. In initial animal trials the drug is described as 'promising and needs to progress to full clinical trials'.

(a) Discuss the ethics of animal testing. [2]

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(b) In a double-blind clinical trial for INX-189, 30 000 patients in several countries were randomly divided into two groups.

(i) Describe what is meant by a *double-blind trial*. [2]

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(ii) Suggest **two** reasons for carrying out the trial with a large number of people from several countries. [2]

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(c) The World Health Organisation monitors the number of cases of different diseases per 100 000 people worldwide. The world's population is 7.5 billion (7.5×10^9). Calculate the number of hepatitis C carriers per 100 000 people. [2]

Number of hepatitis C carriers per 100 000 =

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4. Hospitals started to introduce control measures against MRSA in 2005.

Measles is a disease which was common before 1965, but is now very rare.

Describe the methods used to control the spread of MRSA and measles. Account for the differences in these methods. [6 QER]

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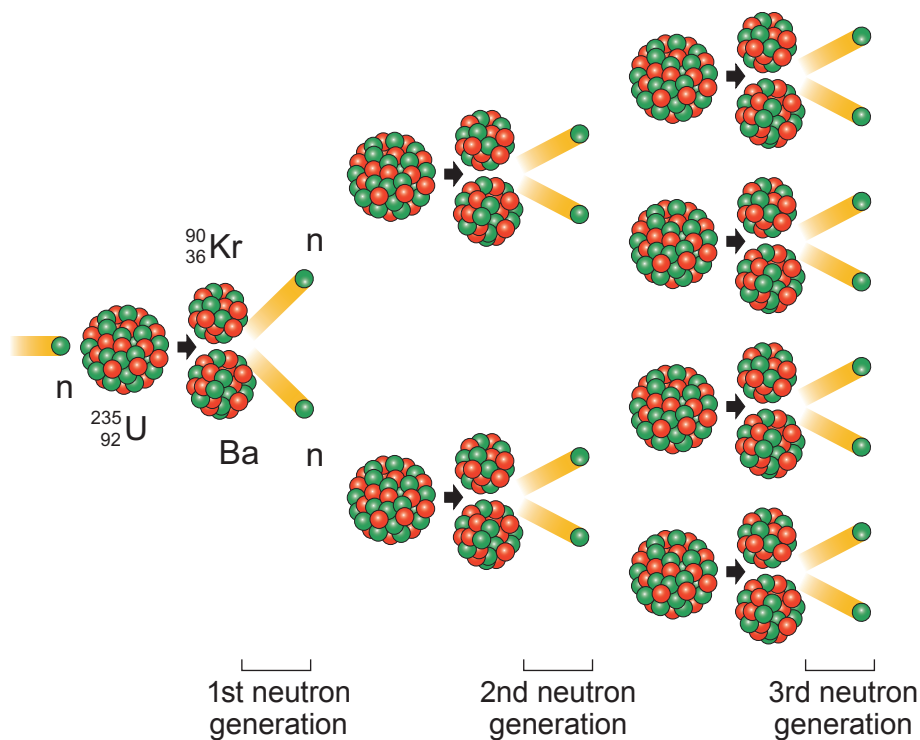
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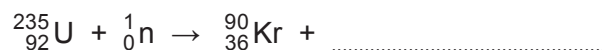
5. Nuclear fission reactors are designed with safety features to prevent nuclear accidents. However in Chernobyl in 1986, during testing, an accident caused a meltdown of the nuclear reactor.

The diagram below shows one possible chain reaction for uranium-235.



(a) Use the information in the diagram to answer the following questions.

- (i) Complete the balanced nuclear equation for the fission reaction shown in the diagram. [2]



- (ii) Complete the table below. [1]

Neutron generation	Number of neutrons released
1 st	2
2 nd	4
3 rd	8
4 th	16
5 th	32
10 th

(b) Explain how an uncontrolled chain reaction could result in a reactor meltdown. [2]

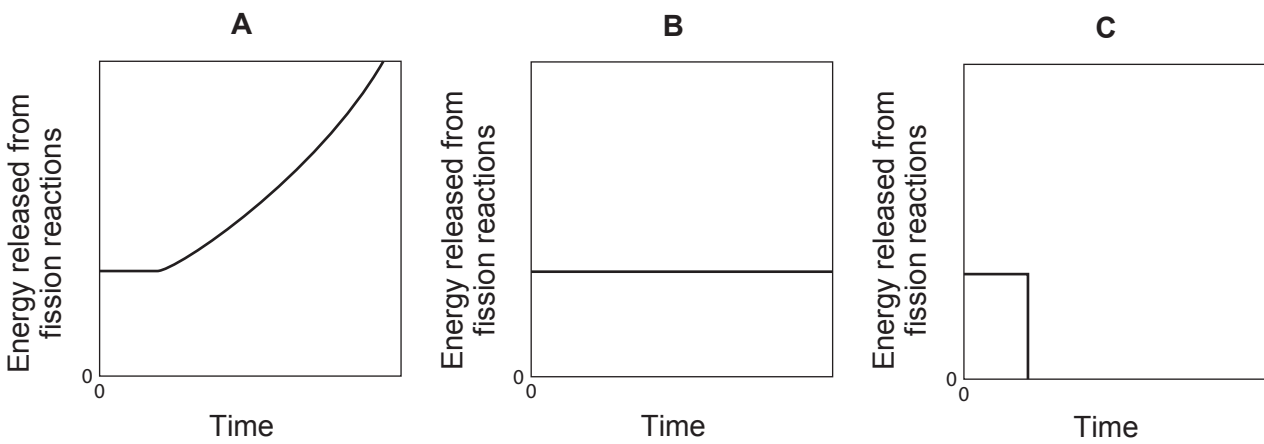
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(c) The graphs below show how the energy released in a nuclear reactor changes when the control rods are in 3 different positions.



Explain the difference in energy production in graphs **A**, **B** and **C** in terms of the relative positions of the control rods. [5]

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6. Doctors can use radioisotopes to treat lung cancer by internal radiotherapy and external radiotherapy.

External radiotherapy uses a radioisotope outside the body. The radiation it emits is targeted at the tumour.

During internal radiotherapy a radioisotope is inserted into the tumour.

- (a) Strontium-90 is a beta source which has a half-life of 29 years. Explain whether you would recommend this radioisotope for use in:

- (i) external radiotherapy. [2]

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- (ii) internal radiotherapy. [2]

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(b) The dose of radiation received for each type of radiotherapy is shown below.

Region	External dose (units)	Internal dose (units)
tissue surrounding tumour	12	4
tumour	20	20

The following results were obtained using the two methods on one type of lung cancer.

	External	Internal
% of patients who survived for at least 5 years after treatment	60	61
% of patients who developed other tumours in the lung within 5 years of treatment	52	35

A patient was diagnosed with a lung tumour. Use the data in the tables to compare the effectiveness of both types of radiotherapy to treat the tumour. [4]

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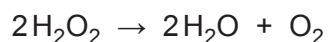
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7. A technician carries out an investigation to find how the rate of decomposition of hydrogen peroxide (H_2O_2) changes with temperature.

She controlled the following variables:

- volume and concentration of hydrogen peroxide
- mass of lead(IV) oxide catalyst

The equation for this reaction is shown below.



- (a) The volume of oxygen gas produced at 40°C is shown in the table below.

Time (s)	Volume of oxygen produced (cm^3)
0	0.0
10	4.4
20	7.2
30	7.6
40	7.6
60	7.6

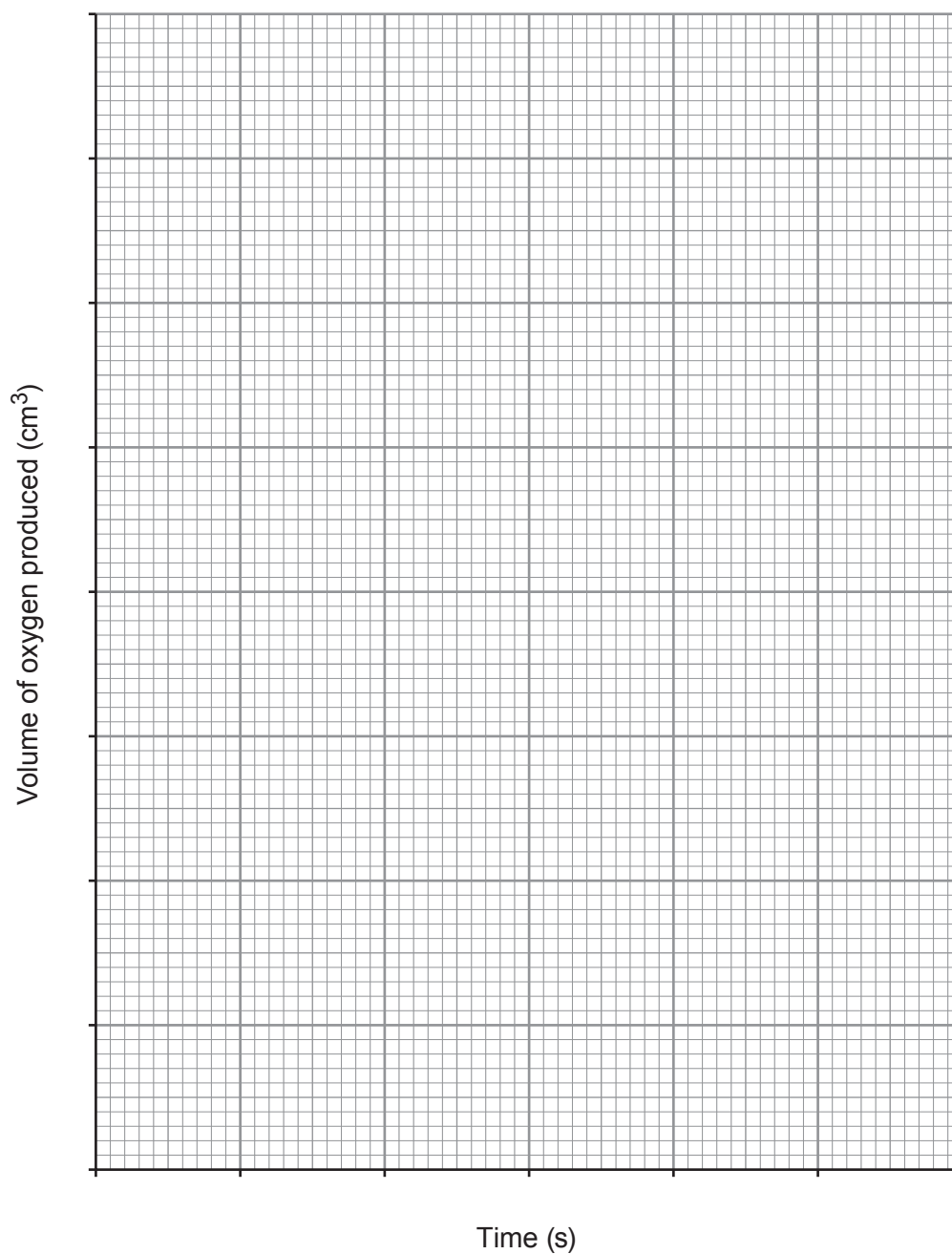
- (i) Plot the data on the grid opposite and draw a suitable line. [3]
- (ii) Add a line to the grid to show the results you would expect if the experiment was repeated at 50°C . [2]
- (iii) Explain in terms of particles, how the rate of reaction is affected by temperature. [2]

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(b) Explain how a catalyst affects the rate of reaction.

[2]

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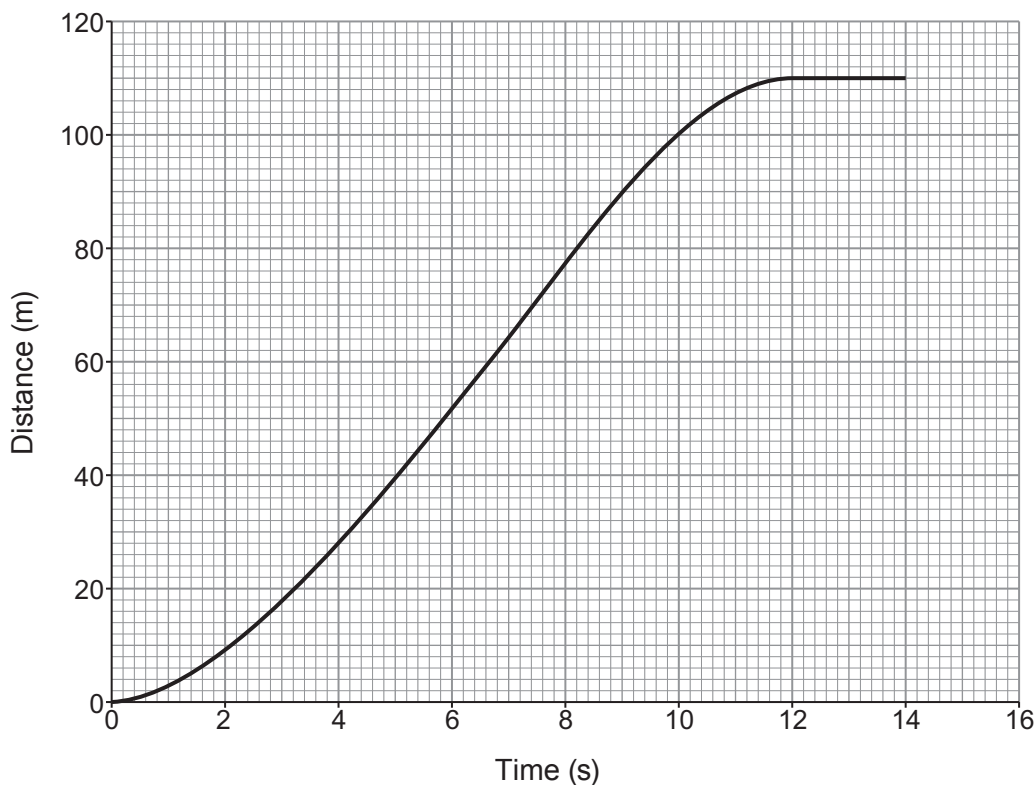
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8. An athlete ran a 100m race in a time of 10 seconds. The distance-time graph for the race is shown below.



- (a) (i) A commentator suggests that the maximum speed of the athlete is 10.0 m/s. Explain whether you agree with the commentator. Add to the graph and show any calculations used in your answer. [4]

Maximum speed =

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- (ii) State why your calculated maximum speed and the mean speed are different. [1]

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(b) Use the information from the graph to explain how the acceleration of the athlete changes during the 14 seconds shown. [4]

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THE PERIODIC TABLE

Group **1** **2** **3** **4** **5** **6** **7** **0**

<div style="border: 1px solid black; padding: 5px; display: inline-block;"> ¹ H Hydrogen ₁ </div>														<div style="border: 1px solid black; padding: 5px; display: inline-block;"> ⁴ He Helium ₂ </div>																		
7 Li Lithium 3	9 Be Beryllium 4			11 Na Sodium 11	12 C Carbon 6	13 Al Aluminium 13	14 Si Silicon 14	15 P Phosphorus 15	16 S Sulfur 16	17 Cl Chlorine 17	18 Ar Argon 18	19 F Fluorine 9	20 Ne Neon 10																			
23 Na Sodium 11	24 Mg Magnesium 12			27 Al Aluminium 13	28 Si Silicon 14	29 Co Cobalt 27	30 Zn Zinc 30	31 P Phosphorus 15	32 S Sulfur 16	33 Cl Chlorine 17	34 Ar Argon 18	35 Br Bromine 35	36 Kr Krypton 36																			
39 K Potassium 19	40 Ca Calcium 20			41 V Vanadium 23	42 Cr Chromium 24	43 Mn Manganese 25	44 Fe Iron 26	45 Ni Nickel 28	46 Cu Copper 29	47 Zn Zinc 30	48 Ga Gallium 31	49 Ge Germanium 32	50 As Arsenic 33	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54															
86 Rb Rubidium 37	87 Sr Strontium 38			89 Y Yttrium 39	90 Zr Zirconium 40	91 Nb Niobium 41	92 Mo Molybdenum 42	93 Tc Technetium 43	94 Ru Ruthenium 44	95 Rh Rhodium 45	96 Pd Palladium 46	97 Ag Silver 47	98 Cd Cadmium 48	99 In Indium 49	100 Sn Tin 50	101 Sb Antimony 51	102 Te Tellurium 52	103 I Iodine 53	104 Xe Xenon 54													
133 Cs Caesium 55	137 Ba Barium 56			139 La Lanthanum 57	140 Ce Cerium 58	141 Pr Praseodymium 59	142 Nd Neodymium 60	143 Pm Promethium 61	144 Sm Samarium 62	145 Eu Europium 63	146 Gd Gadolinium 64	147 Tb Terbium 65	148 Dy Dysprosium 66	149 Ho Holmium 67	150 Er Erbium 68	151 Tm Thulium 69	152 Yb Ytterbium 70	153 Lu Lutetium 71	154 Hf Hafnium 72	155 Ta Tantalum 73	156 W Tungsten 74	157 Re Rhenium 75	158 Os Osmium 76	159 Pt Platinum 78	160 Au Gold 79	161 Hg Mercury 80	162 Tl Thallium 81	163 Pb Lead 82	164 Bi Bismuth 83	165 Po Polonium 84	166 At Astatine 85	167 Rn Radon 86
223 Fr Francium 87	226 Ra Radium 88			227 Ac Actinium 89																												

Key

A _r	Symbol Name	Z
Z		

relative atomic mass

atomic number